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Practice:

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questions. This is the  
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## **Electrochemistry questions (practice)**

**| Khan Academy**

Fuel Cells & Galvanic  
*Page 5/26*

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Cells Chapter Exam

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identifies the cathode  
and gives the "a p  
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## **Galvanic Cell Sample Exam Questions With Answers**

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At 25° C, the standard cell potential for this galvanic cell in volts is  
a. 0.13 d. 0.21 b. 0.34  
e. 0.94 c. 0.47 \_\_\_\_ 15.

This question consists of an assertion (statement) in the left-hand column and a reason in the right-hand column. Assertion Reason In the electrochemical cell: the zinc electrode is the negative terminal of the cell.



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**Electrochemistry-  
Ch.9 - PRACTICE  
TEST/EXAM  
QUESTIONS**

e. A spontaneous electrochemical process has electromotive force:  $E > 0$   
f. A spontaneous electrochemical cell is called: A Nernst cell  
An Electrolytic cell  
A Galvanic cell  
g. A cell using electricity to enable a non-spontaneous process is

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called: A Nernst cell An  
Electrolytic cell A  
Galvanic cell h.

## **Electrochemistry Exercises**

Electrochemistry &  
Voltaic Cells Chapter  
Exam Instructions.  
Choose your answers  
to the questions and  
click 'Next' to see the  
next set of questions.

## **Electrochemistry & Voltaic Cells - Practice Test**

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(c) Reduction occurs at the zinc electrode as the cell operates. (d)

The mass of the zinc electrode will decrease as the cell operates. (e)

The concentration of  $\text{Ag}^+$  will decrease as the cell operates. 13.

Consider the standard voltaic (or galvanic)

cell:  $\text{Fe}, \text{Fe}^{2+}$  versus  $\text{Au}, \text{Au}^{3+}$ . Which

answer identifies the cathode and gives the

$E^\circ$  for the cell?

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## **Sample Questions - Chapter 21**

Each sample has a mass of 3.20 g. i.

Sample 1 is placed in the calorimeter and completely reacted with oxygen. The temperature rise of the calorimeter is found to be  $2.21^{\circ}\text{C}$ . Calculate the heat of combustion of the fresh lignite in  $\text{kJ g}^{-1}$ . 2 marks ii.

Sample 2 is placed in an oven for 2 hours

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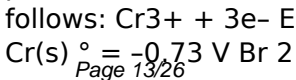
where it is kept at a temperature of 100°C.

## **SAMPLE EXAM QUESTIONS**

Use the following to answer questions 9-10:

Consider the galvanic cell shown below (the contents of each half-cell are written beneath each

compartment). The standard reduction potentials are as follows:



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Questions With  
Answers

(aq) + 2e<sup>-</sup> - 2Br<sup>-</sup> E° =  
+1.09 V 9. What is E°  
for this cell? A) 1.82 V  
B) 0.36 V C) 4.75 V

## Ch 11 Practice Problems - UCSB

A.P. Chemistry Practice  
Test - Ch. 17:

Electrochemistry ... The  
purpose of the salt  
bridge in an  
electrochemical cell is  
to \_\_\_\_\_. A) provide a  
source of ions to react  
at the anode and  
cathode. B) provide

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oxygen to facilitate  
oxidation at the anode.  
... questions 2&3 in  
previous years.

## **A.P. Chemistry Practice Test - Ch. 17: Electrochemistry A ...**

Practice Questions.  
Khan Academy.  
Studying metabolism  
with galvanic cells.  
MCAT Official Prep  
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Online Flashcards  
Question 17, Chemistry

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Question Pack Passage

19 Question 111.

Sample Exam  
Practice Exam 2 C/P

Questions With  
Section Passage 6

Answers  
Question 29 Key Points

- An electrochemical cell, also called galvanic or voltaic cell, is a device that produces ...

## **Galvanic Or Voltaic Cells -**

## **Electrochemistry - MCAT Content**

These are actual  
Provincial Exam



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questions! Your own  
provincial exam and  
unit test will ... In an  
operating  
electrochemical cell

the function of a salt  
bridge is to . ... Which  
of the following metals  
could be used to  
cathodically protect a  
sample of lead? A. iron  
B. gold C. silver D.  
copper . 104. V0 4 .

SAHOTA 03

Electrochemistry Study  
Guide ...

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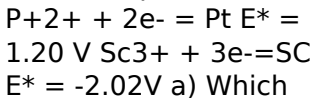
Galvanic Cell  
Chemistry questions in  
Sample Exam  
detail.

Questions Withe  
Answers

## Questions and Answers of

### Electrochemistry Electrochemical ...

Answer the following  
questions based on a  
galvanic cell at 25°C  
and the following  
electrode potentials:



a) Which  
chemical species are  
the oxidant and

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Answers

reductant? Oxidant:  
Reductant: b) Does  
oxidation or reduction  
happen at the  
cathode?

## **Solved: Answer The Following Questions Based On A Galvanic**

...

CliffsNotes study  
guides are written by  
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can ease your

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homework headaches  
and help you score  
high on exams.

## **Quiz:** **Electrochemical Cells**

A basic run-through of  
a typical HSC question  
on galvanic cells.

Question reference:  
2008 Board of Studies  
HSC Chemistry Exam,  
question 25. **\*\*NOTE:**  
This is a SAMPLE video.  
If you would like to see

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## **HSC Chemistry Exam Question - Galvanic Cells**

A galvanic cell (which is also sometimes referred to as a voltaic or wet cell) consists of two half-cells, which convert chemical potential energy into electrical potential energy. Galvanic cell A galvanic cell is an electrochemical cell which converts chemical potential

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energy to electrical  
potential energy  
through a spontaneous  
chemical reaction.

## **Galvanic And Electrolytic Cells | Electrochemical ...**

Because all of these  
questions relate to the  
same cell, I have  
included all of them  
together in this section  
as drbp down boxes so  
you don't have to jump  
around as much. Each  
answer is worth 1

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point, just like on a paper exam. A standard galvanic cell is made from two half reactions shown below.

## **Solved: Because All Of These Questions Relate To The Same**

...

Test Of Electrochemical Cell Experiment 10

Questions | By

Pipitcupit | Last

updated: Aug 31, 2017

| Total Attempts: 166

Questions All questions



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5 questions 6

questions 7 questions

8 questions 9

questions 10 questions

Answers

**Test Of**

**Electrochemical Cell**

**Experiment -**

**ProProfs Quiz**

Galvanic Cell or Voltaic

Cell - A Galvanic cell,

also known as the

Voltaic cell is a device

in which electrical

current is generated by

a spontaneous redox

reaction. A galvanic

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cell has two half cells.  
Click here for detailed  
content on the  
Galvanic cell (Voltaic  
Cell).

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ecf8427e.