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Chapter 15 Water and Aqueous Systems159 SECTION 15.1 WATER AND ITS PROPERTIES (pages 445–449) This section describes the properties of water in the liquid and solid states and

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explains how hydrogen bonding affects the surface tension and vapor pressure of water. Water in the Liquid State (pages 445-447)

## **SECTION 15.1 WATER AND ITS PROPERTIES (pages 445-449)**

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of water in terms of the structure of the  
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Chapter 15. Solution Dynamics. 573. or a chemical reaction to take place, the reacting particles must first collide—freely, frequently, and at high velocity. Particles in gases move freely and swiftly and collide continually, but gas phase reactions are difficult to contain and are sometimes dangerous.

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Lecture Notes for Chapter 15: Solutions & Solution Chemistry. I. Solutions a. A . solution. is simply a homogeneous mixture i. Homogeneous: same

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throughout (it does not mean “one”) ex:  
water + sugar, air, alloys, aerated water  
ii. Heterogeneous means different  
throughout (does not mean “two”) ex:  
oil + water b. Solute

## **Lecture Notes for Chapter 15: Solutions & Solution Chemistry**

Chemistry Workbook Chapter 15 Water  
And Aqueous Systems ... In the world of  
chemistry, an aqueous solution is any  
solution that contains water as the  
solvent. A solution is a mixture of two or  
more substances made of a solute,  
which dissolves in the solvent. A liquid,  
on the other hand, consists of molecules  
or atoms with connecting

## **Chemistry Workbook Water And Aqueous Systems Answers**

Chapter 15 – Solutions 15.1 Definitions  
related to Solutions □Solution- uniform  
mixture of two or more substances. A  
solution is composed of a solute(s)  
dissolved in a solvent. □Solute- the  
substance present in lesser amount in a

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solution. □ Solvent- the substance present in greatest amount in a solution.

## **Chapter 15 Solutions**

Chapter 15 Solutions. 1. A homogeneous mixture is a combination of two (or more) pure substances that is uniform in composition and appearance throughout. Examples of homogeneous mixtures in the real world include rubbing alcohol (70% isopropyl alcohol, 30% water) and gasoline (a mixture of hydrocarbons). 2.

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Depiction in chemical reactions. In this chapter, we will learn about some



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reactions in which the key steps involve the movement of single electrons. You may recall from way back in section 6.1A that single electron movement is depicted by a single-barbed 'fish-hook' arrow (as opposed to the familiar double-barbed arrows that we have been using throughout the book to show two-electron movement).

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15.8 Biologically Important Molecules

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(15.8A) 15-25 Waxes Fats and Oils  
Amides (15.8B) 15-27 Proteins Peptides  
Lactams (15.8C) 15-28 15.9  
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