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(Chapter 13) Multiple Choice (20%) 1)

Please write the equilibrium expression for this reaction. $3 A (aq) + 3 D (s) \leftrightarrow 4 C (l) + B (g) + 2 E (s)$ 2) Please consider

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the gas phase reaction: $2 \text{NO} + \text{Br}_2 \leftrightarrow 2 \text{BrNO}$ $\Delta H = -345 \text{ kJ/mol}$. Which one would increase the yield of NO?

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13...Take 2)

A.P. Chemistry Practice Test - Ch. 13:
Equilibrium MULTIPLE CHOICE. Choose
the one alternative that best completes
the statement or answers the question.

B)the rate constants of the forward and
reverse reactions are equal. A.P.

Chemistry Practice Test - Ch. 13:
Equilibrium

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Name _____

MULTIPLE CHOICE.

A.P. Chemistry Practice Test - Ch. 13: Equilibrium ...

AP Chemistry: Chapter 13--Equilibrium
Chapter 13 Test Review Know how to write K_c , K_p expressions (*note that solids and liquids not included; K_c in [],

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and K_p in atm) Know how to calculate K_c , K_p when: Initial concentrations (or pressures) are known Changes are known Equilibrium concentrations (or pressures) are known *Note for each of the above, ICE diagrams are very helpful!!!

AP Chemistry: Chapter

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13--Equilibrium Chapter 13 Test ...

You are given a box containing NH_3 , N_2 , and H_2 at equilibrium at 1000°C .

Analysis of the contents shows that the concentration of NH_3 is 0.102 mole/liter, N_2 is 1.03 moles/liter, and H_2 is 1.62 moles/liter. Calculate K for the reaction: $2\text{NH}_3(\text{g}) \leftrightarrow \text{N}_2(\text{g}) + 3\text{H}_2(\text{g})$

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AP Chemistry Review Questions - Chemical Equilibrium

Mr. Mac's AP Chemistry Site Ken
MacGillivray - Hoggard High School.
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» **Chapter 13 - ap.kmacgill.com**

Chapter 13 - Chemical Equilibrium . Intro
. A. Chemical Equilibrium 1. The state

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where the concentrations of all reactants and products remain constant with time

2. All reactions carried out in a closed vessel will reach equilibrium a. If little product is formed, equilibrium lies far to the left b.

Chapter 13 - Chemical Equilibrium - ScienceGeek.net

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AP Chemistry; Ch 1 and 2: Scientific Notation and Unit Analysis. Matter Handout. ... Ch 13: Equilibrium Handout Key . Equilibrium Practice Answers . AP Question . Ch 14 and 19: Ch 14 Ch19 Practice Keys . Ch 15 pH Calculation Answers. Chapter 15 Acid Base Practice Test and ACS Answers. Ch 15 Handout Keys . Ch 16 Titration

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Curves/Calculations Answers

Baker, Mrs. (Science) / AP Chemistry

Designing experiments and procedures to test a prediction or theory. Creating graphs, diagrams, and models that represent chemical phenomena.

Explaining how the microscopic structure of a substance determines its

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chemical properties. ... AP Chemistry Course and Exam Description

AP Chemistry - AP Students | College Board

The Organic Chemistry Tutor 236,224 views. ... AP Chemistry Introduction to Chemical Equilibrium - Duration: ... Chapter 13 - Properties of Solutions: ...

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Chapter 13 (Chemical Equilibrium) - Part 2

AP Chemistry Contact AP Chemistry - Semester 1. Semester 1 Test Review. Days 1-10. Assignments. Stoichiometry Review. Test Review. ... Key - Chapter 13 and 14 Test Review. KEY - Chapter 13/14 Multiple Choice Rev. Due Date Lab Due

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12/12/18. Chapter 15 - Equilibrium .
Assignments .

AP - SEMESTER 1 - roosevelt high school

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This test covers the essentials of Mitosis and Meiosis cell divisions and is helpful

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test.

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questions and all short answer must be justified with an adequate explanation or show clearly all calculations leading up to the final answer. ... Unit 9 Chapter 13 (1 week) Solutions.

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Chapter 13 - 14 Practice Quiz

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